

WRITING NET IONIC EQUATIONS: Practice

CHEMISTRY BC2001x

Write *balanced net ionic equations* describing the predominant reaction(s) that occur when the substances below are mixed in aqueous solution. None involves oxidation-reduction. If nothing happens, write “no reaction”. Check your work on the web.

Zinc hydroxide reacts with excess acetic acid

Potassium iodide solution is mixed with mercury (I) nitrate solution

Solid barium sulfite is mixed with aqueous nitric acid

Solutions of lead acetate and ammonium sulfate are mixed.

Barium nitrate solution is mixed with sodium iodate solution

Sodium hydroxide solution is mixed with nitrous acid

Hydrochloric acid is mixed with potassium nitrite solution

Hydrochloric acid is mixed with potassium hydroxide

Silver acetate solution is mixed with hydrochloric acid

Sodium acetate solution is mixed with nitric acid

Ammonium chloride solution is mixed with sodium carbonate solution

Copper II hydroxide is mixed with acetic acid.

Lead hydroxide is mixed with hydroiodic acid.

Solutions of iron (III) nitrate and sodium sulfide are mixed

Solutions of potassium phosphate and aluminum chlorate are mixed.

A solution of potassium phosphate is mixed with excess nitric acid.

A solution of zinc chlorate is mixed with perchloric acid.

Solutions of zinc chlorate and sodium hydroxide are mixed.