

SOLUBILITY/ELECTROLYTE PRACTICE

Chemistry BC2001x

Applying rules on solubilities and strong/weak electrolytes

Fill in all the blank spaces, following the form used in the example given.

Chemical Substance	Chemical Formula	Solubility in Water	Strong / Weak Electrolyte	Reason(s) Rules Applied or Exceptions Noted
sodium chloride	NaCl	soluble	strong	Na is an alkali metal, and all of the common alkali metal salts are soluble. Almost all salts are strong electrolytes.
hydrogen sulfide = hydrosulfuric acid				
barium sulfate				
ammonia				
calcium chloride				
lithium nitrate				
potassium sulfide				
lead carbonate				
strontium sulfate				
nickel perchlorate				
acetic acid				

Chemical Substance	Chemical Formula	Solubility in Water	Strong / Weak Electrolyte	Reason(s) Rules Applied or Exceptions Noted
mercurous iodide = Hg(I) iodide				
ammonium chloride				
sodium acetate				
nitric acid				
lead sulfate				
silver bromide				
cupric nitrate = copper (II) nitrate				
silver carbonate				
potassium carbonate				
zinc sulfide				
lithium hydroxide				
hydrogen cyanide = hydrocyanic acid				